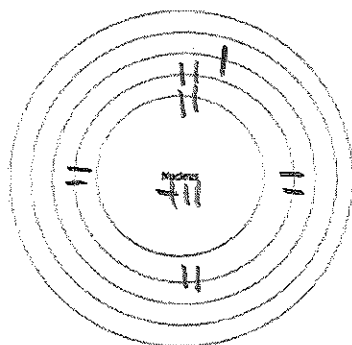


Bohr Model and Valence Electron Practice

Directions: For each atom, draw the Bohr model. The circles represent the electron shells. Then draw the Lewis dot diagram and state if it will gain or lose electrons and by how many to become stable or noble gas.

Sodium

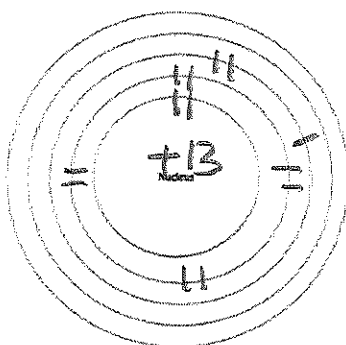


Lewis Dot Diagram



Gain ___ e⁻ Lose 1 e⁻
Stable

Aluminum

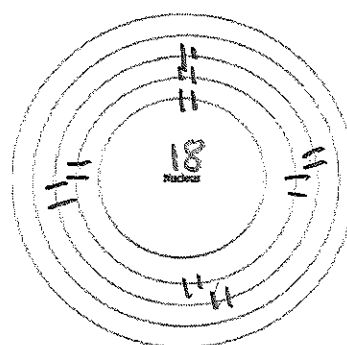


Lewis Dot Diagram



Gain ___ e⁻ Lose 3 e⁻
Stable

Argon

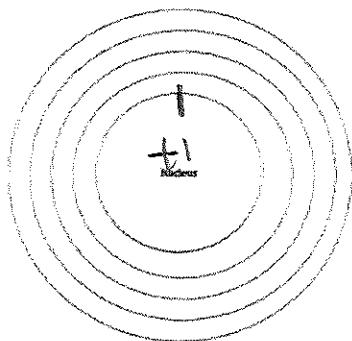


Lewis Dot Diagram

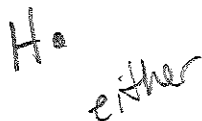


Gain 0 e⁻ Lose 0 e⁻
Stable

Hydrogen

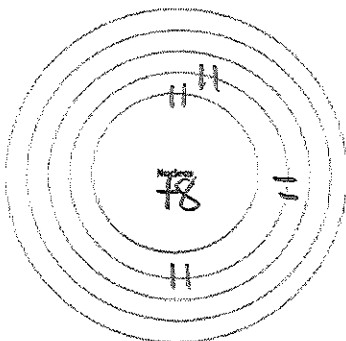


Lewis Dot Diagram



Gain 1 e⁻ Lose 1 e⁻
Stable

Oxygen

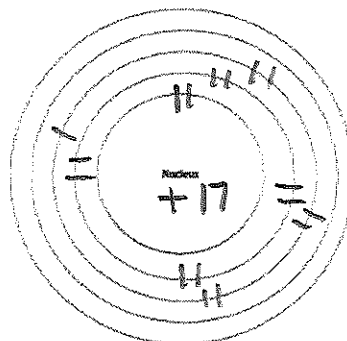


Lewis Dot Diagram

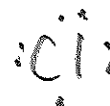


Gain 2 e⁻ Lose ___ e⁻
Stable

Chlorine



Lewis Dot Diagram



Gain 1 e⁻ Lose ___ e⁻
Stable

Go back and highlight the valence electrons in the Bohr model.